

Application of Electrodialysis in Waste Water Treatment and Impact of Fouling on Process Performance

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Abstract

Electrodialysis (ED) is a new advanced separation process that is commonly utilized for producing drinking water from water bodies as well as for the treatment of industrial effluents. ED process is applied on commercial scale. Basically, an ED process consists of an ion exchange membrane and the driving force necessary for applicability of the process is electric potential. Due to the presence of electric potential ions from one solution after passing through ion selective membrane barrier are transferred to another solution. The main factors on which ED process performance depends on concentration of ion in raw water, flow rate, concentration of feed, current density, membrane properties and cell compartments geometry.

Fouling which is produced by foulants including organics, colloids and biomass on the inside membrane internal structure or on the outside surface results in reduction of process separation efficiency and energy consumption is enhanced. Fouling increases the membrane resistance and selectivity of membrane is reduced by fouling. Therefore, some methods are proposed to reduce fouling in ED system such as pre-treatment of feed solution, zeta potential control, membrane properties modification and flowrate optimization. It is a need of an hour to suggest a reducing method less energy and thus minimum operating and investment cost. Electrodialysis Reversal (EDR) system can be as regarded as best option because no extra chemicals are required and life of membrane increased by it. In EDR fouling progress is broken by revering electric potential (applied electric field). This paper elaborates ED process briefly and presents an overview of literature review on different kinds of fouling mechanisms. Also, different cleaning methods have been briefly described for enhancing efficiency of ED process.

Keywords: Wastewater; Fouling; Permselectivity; Ion exchange membrane

Abbreviations: ED: Electrodialysis; EDR: Electrodialysis reversal; TDS: Total Dissolved Solid; BPM: Bipolar Membrane

Introduction

The population of world is increasing day by day and natural water resources have been found insufficient to meet the demand of fresh water. Water scarcity could be the global risk for next decades if steps will not be taken for the reclamation of waste water. Method of waste water treatment is chemical coagulation, biological degradation, electrochemical methods, adsorption and membrane. Among these techniques membrane technology has been found significant because of treated water quality. Membrane is defined as a selective barrier between two phases and it can be formed from natural and synthetic material including organic and inorganic polymer, ceramic and metal material. Separation depends upon pore size [1].

ED process is basically a membrane separation process and it is commonly used for the treatment of waste water (Figure 1) [2]. In ED electric potential is used as a driving force and ion exchange membrane is applied between anode and cathode. On the application of electric current negative and positive ions are moved towards the respective electrodes based on their polarity. Cations are passed through the compartment containing membrane having negative charge and vice versa for anions. The compartments are diluted and concentrated in alternative way. The performance of an ED process depends upon following parameters current density, pH, flow rate, ED cell structure, feed water ionic concentration, properties of ion exchange membrane. Membrane fouling is an important factor that results in enhancement of energy consumption and decline of membrane flux. Selectivity of membrane is adversely reduced. Fouling is caused by the deposition

of biomass, colloid and organics substances. Accumulation could be happened on the inside or on the outside surface of membrane. As most of collides presents in water are negatively charged so most of the research is carried out on anion exchange membrane fouling [3].

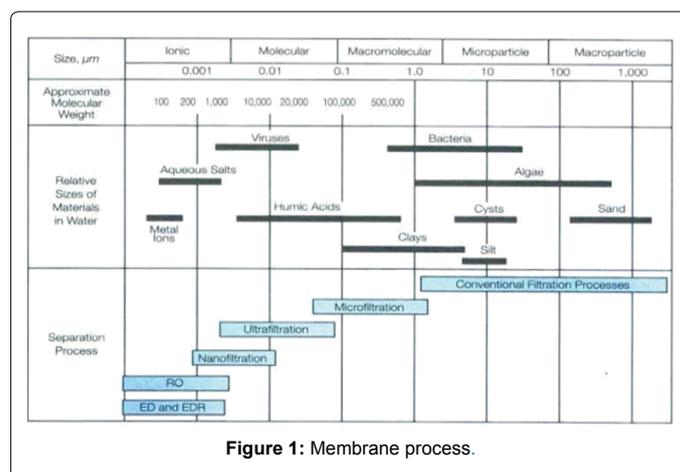


Figure 1: Membrane process.

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Working principle of ED system

The main objective of ED process is removal of salts from aqueous solution by passing an aqueous solution through an ion exchange membrane (Figure 2) [4]. The cell is divided into compartments by placing anion and cation exchange membrane between two electrodes acting a cathode and anode. Direct current is applied between two electrodes. Consider feed solution containing sodium chloride passes through ED system and focus on one compartment. Cations pass through cation exchange membrane and anions are restricted. Anion could only pass through anion exchange membrane which restricts passage of cations. The concentration of salts decreases in this compartment but salt concentration keeps increasing in the adjacent compartment. Thus, salt concentration decreases in alternative compartments and enhances in the rest. Both desalinated water and concentrated brine leave from the adjacent compartments. Applied electric potential acts as driving force for ED process.

ED is preferred when salt concentration is low and it regarded as an energy efficient process. The electric current requirement depends upon number of ions transported through ion exchange membrane.

$$\text{Electric current required} = Z \times F \times Q \times \Delta C / \epsilon$$

Where, Z=Charge on ion; F=Faraday's constant; Q=Feed solution flow rate; ΔC =Change in solute concentration; ϵ =Current utilization factor that accounts for energy efficiency.

ED processes have been used widely on commercial scale and various applications of ED systems have been shown in Table 1 [5].

ED process was first time used by Maigrot and Sabates for demineralizing sugar syrup in 1890 (Figure 3) [6]. Wood was used to manufacture frame of ED system. Anode and cathode were formed by carbon and permanganate paper was used to make ion exchange membrane. In order to provide electric current dynamo was used. The syrup solution was poured through anodic compartment and anode was placed in middle of the system. On the application of electric potential cations such as sodium, magnesium and calcium were transferred

through anodic region to cathodic region. Demineralized syrup was obtained in anodic compartment [6].

Removal of copper ions by the application of ED system was study by [3]. In order to achieve copper ions removal two different ion exchange membranes were used. Following parameters were determined initial copper ions concentration, pH, flow rate, conductivity of solution. At the value of 7.5 V applied voltage and concentration of copper initially present 4 g/L the percentage of copper ion removal greater than 85% was obtained. The effect of temperature, voltage and flow rate on the performance of ED system used for lead removal was studied [7]. It was reported that cell performance enhanced by increasing voltage and temperature but decrease in separation percentage was revealed by increasing flow rate.

ED process can be used for the removal dissolved solid having charge on their surface by the application of ion exchange membrane. The membrane used for cations transfer is called cation exchange membrane and it restricts anions. The function of anion exchange membrane is vice versa to cation exchange membrane. The thickness of membrane is 0.1 to 0.6 and based upon morphology membrane could be homogeneous or heterogeneous. The ion exchange membrane is made by making blend of ion exchange resin with polymer and ion exchange membrane must have at least half of polymer mixture. Some of the important features that are essential for cation or anion exchange membrane are ion exchange membrane must have low electrical resistance, it should show insolubility in aqueous solution, membrane must have long life expectancy, it can be used above 46°C, semi rigidity for ease of handling, membrane should be available in different sizes and configurations. 50-99 % removal of source water pollutants or dissolved solids could be achieved by ED process [8]. Total Dissolved Solid (TDS) can be treated and production of water less than 10 mg/L TDS is achieved by ED system when water salinity in the range of 100 to 1200 mg/L. The main advantage of ED process is no phase change takes place, no additional chemical requirement, high perm selectivity, low electrical resistance, high mechanical and chemical stability, some industrial applications of ED process have been depicted in Table 2 [5].

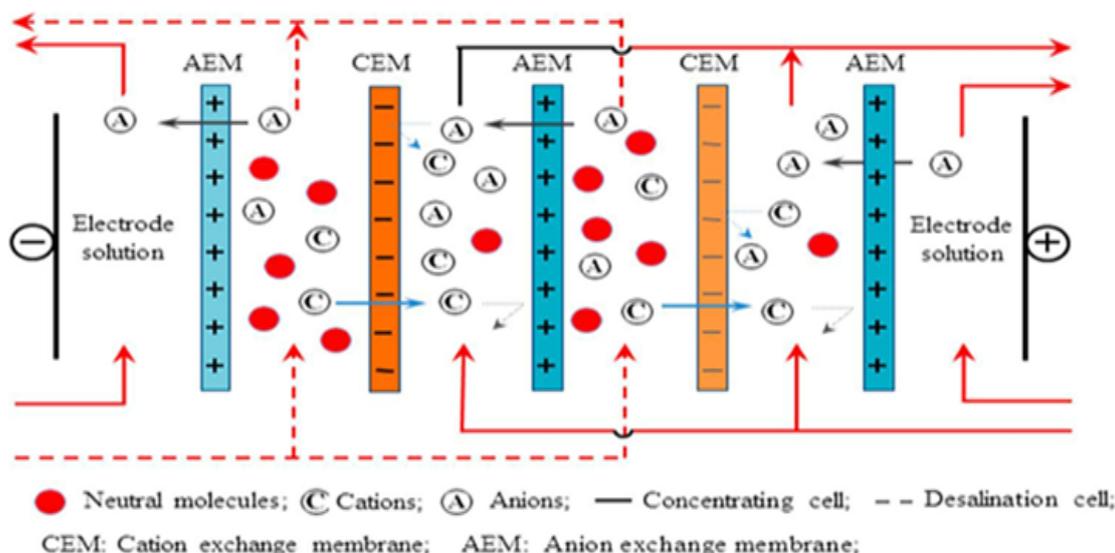


Figure 2: Basic principle of electro dialysis.

Processes	Applications
Electrodialysis	Used for water desalination and salt pre-concentration
Donnan dialysis	For water softening
Reverse ED	For generating electro dialytic energy
Diffusion dialysis	For recovering acid and base
Capacitive deionization	It is used for water desalination and water softening

Table 1: Applications of processes related to electrodialysis system.

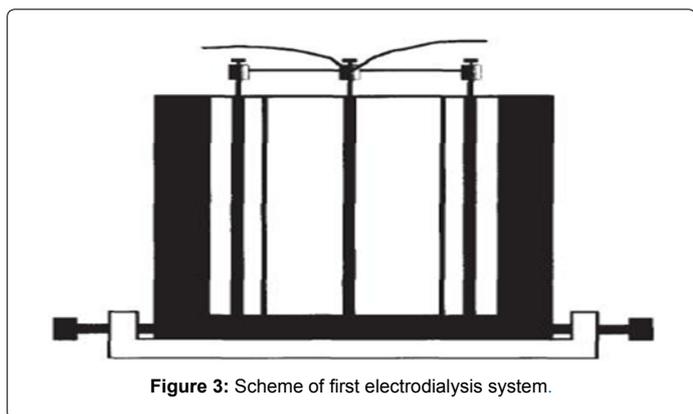


Figure 3: Scheme of first electrodialysis system.

Applications	Scale	Major Problems
Desalination	Applied commercially	Cost
Water treatment process	Applied commercially	Cost
Waste treatment	Applied commercially	Fouling of membrane
Demineralization of food products	Pilot scale plant	Fouling of membrane
Production of table salt	Commercial	Fouling of membrane

Table 2: Industrial applications and major problems of ED.

Ion exchange membranes used for ED process

Depending on the type of ionic groups attached to the membrane matrix, ion exchange membranes are classified into anion exchange membranes and cation exchange membranes.

Negatively charged groups, such as $-SO_3^-$, $-COO^-$, $-PO_3^{2-}$, $-PO_3H^-$, $-C_6H_4O^-$, etc., formed cation exchange membranes which are fixed to the backbone of membrane reject anions while allowing the passage of cation.

Charged groups, such as $-NH_3^+$, $-NRH_2^+$, $-NR_2H^+$, $-NR_3^+$, $-PR_3^+$, $-SR_2^+$, etc., formed anion exchange membranes which are fixed to the backbone of membrane reject cations while allowing the passage of anions (Figures 4 and 5) [9,10].

Homogeneous ion exchange membranes

Different techniques are available to blend ionic groups for the preparation of homogeneous ion exchange membranes. On the basis of starting material membrane, they are classified into 3 types [11,12] but the two most commonly used are,

(a) Anionic or cationic exchange groups monomer to form an ion exchange membrane with the help of copolymerization process;

(b) Use polymer film, introduce ionic characters in two ways directly with functional monomer using grafting process or indirectly monomer non-functional grafting with the help of functionalization reaction (Figure 6) [10].

Inorganic-organic ion exchange membranes

Due to the exceptional properties within single molecule composites inorganic-organic composite materials are very important which enhance the interaction between the two different molecules [13]. In comparison to the refined organic polymer and inorganic materials these materials have extraordinary properties in terms of their mechanical [14], thermal [15], electrical [16], and magnetic properties [17]. Organic polymer shows high photoconductivity, efficient luminescence, structural flexibility, convenient processing and even metallic characteristics. High carrier mobilities, band gap tunability, magnetic field range and dielectric properties and stability in terms of thermal and mechanical which is provided by inorganic compounds. The interface attraction between the organic and inorganic components of different material by combining them in different ways new properties also evolve [18-20].

In recent years due to different application of inorganic-organic material with the combination of different composition the use of inorganic-organic material for ion exchange membrane production is increased. For fuel cell separator application of ion exchange membranes prepared from inorganic-organic material in increased due to the high thermal stability [21,22].

Bipolar ion exchange membranes

Membrane compositions which contain a layer of selective cation (contain fixed negative charges) and a layer of selective anions (contain fixed positive charges) are called Bipolar Membrane (BPM). Separation

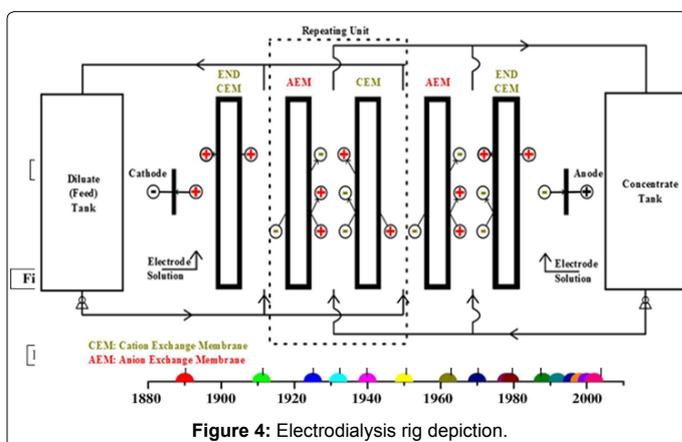


Figure 4: Electrode dialysis rig depiction.

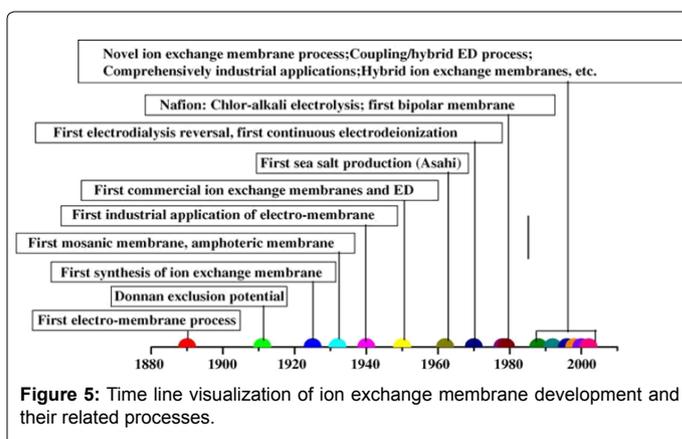


Figure 5: Time line visualization of ion exchange membrane development and their related processes.

of mono and divalent ions, anti-deposition, anti-fouling, water dissociation, etc. novelties come with the help of composition of anionic and cationic exchange layer [23-25].

Production of acids and bases from the corresponding salts especially for electroolytic water splitting with the help of BPMs is shown in Figure 7 [26]. This growth point become a very important in many industrial process and daily life technique [24,27]. Conventional technique which contaminates the environment changes with the help of BPM [28].

Amphoteric ion exchange membranes

Structure in which weak acid (negative charge) and weak basic (positive charge) randomly distributed inside the membrane matrix is called amphoteric ion exchange membranes [29]. Response to the

external solution is given by these membranes in term of pH which shows the charge present on these membranes (Figure 8) [30].

Amphoteric ion exchange membranes were first suggested by Söllner in 1932 together with mosaic membranes. A lot of research work has been conducted for this type of membrane. The description of ions permeation through amphoteric membrane using a specialized amphoteric membrane model was proposed by Takagi and Nakagaki for describing permeation of ions through and also, they determined charge by amphoteric membrane dissociation.

Permselectivity of cations through membrane having negative charge on the surface

Permselectivity is an important parameter for evaluating performance of membrane and for cations exchange membrane that

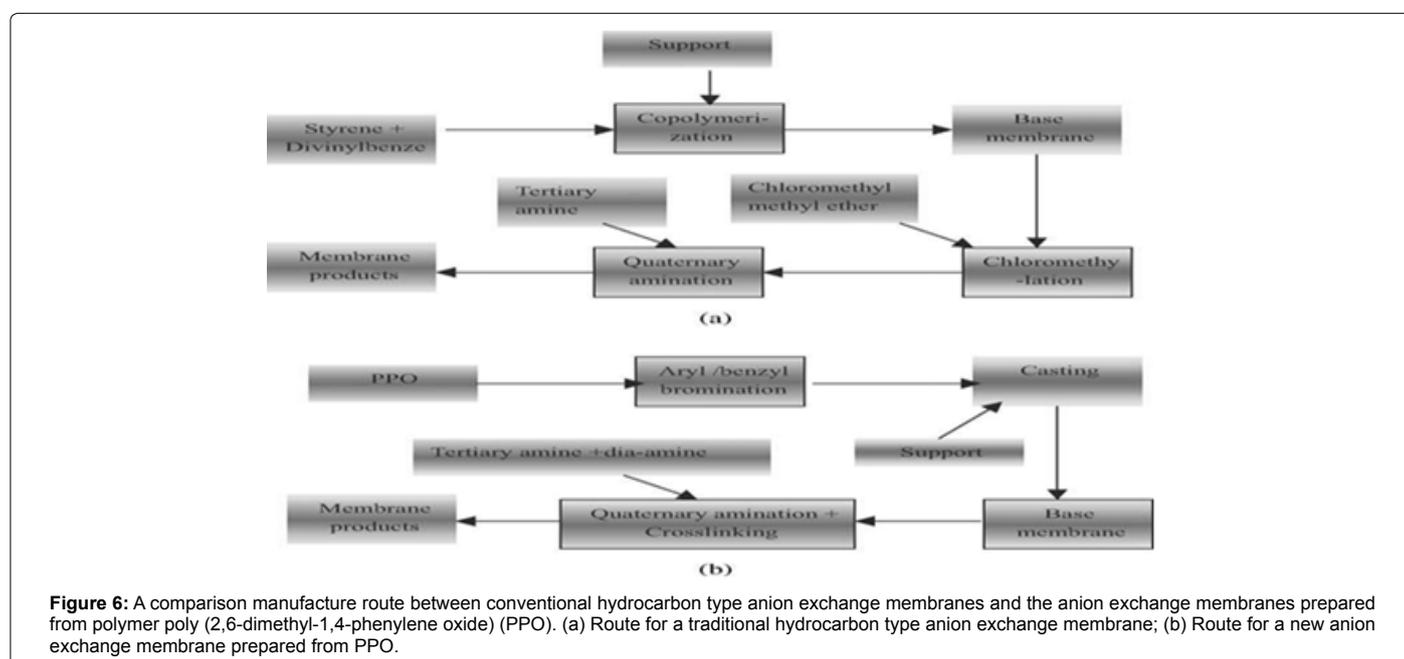


Figure 6: A comparison manufacture route between conventional hydrocarbon type anion exchange membranes and the anion exchange membranes prepared from polymer poly (2,6-dimethyl-1,4-phenylene oxide) (PPO). (a) Route for a traditional hydrocarbon type anion exchange membrane; (b) Route for a new anion exchange membrane prepared from PPO.

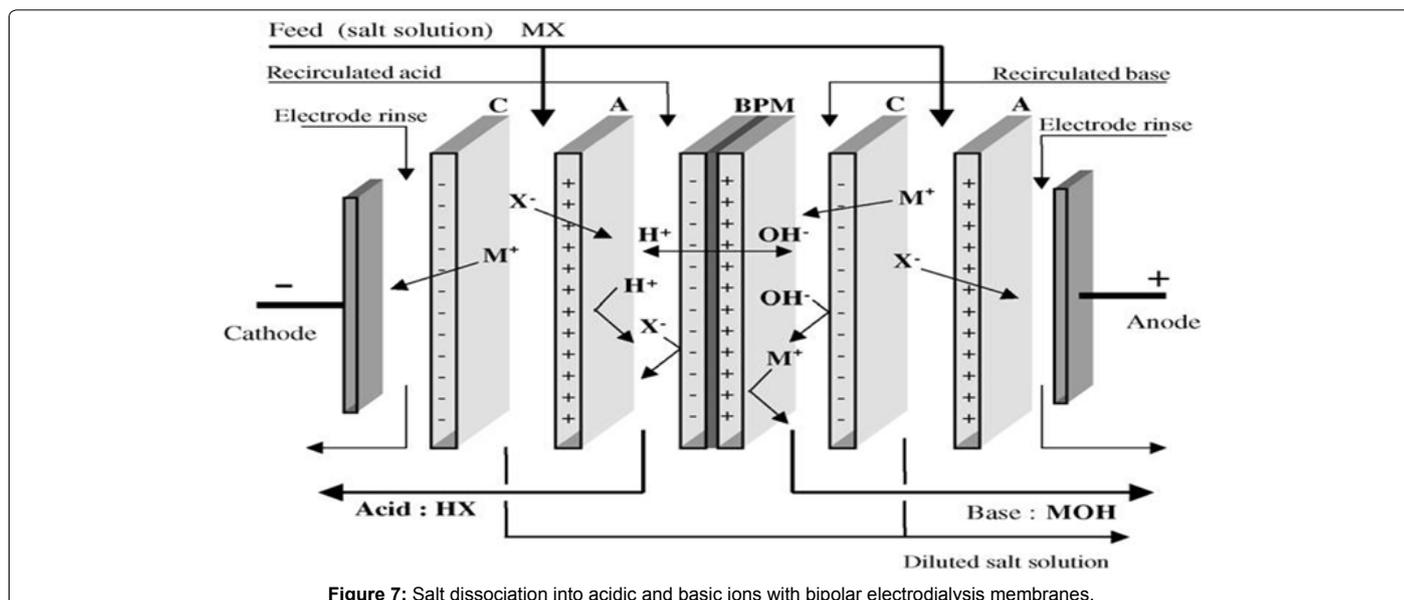


Figure 7: Salt dissociation into acidic and basic ions with bipolar electro dialysis membranes.

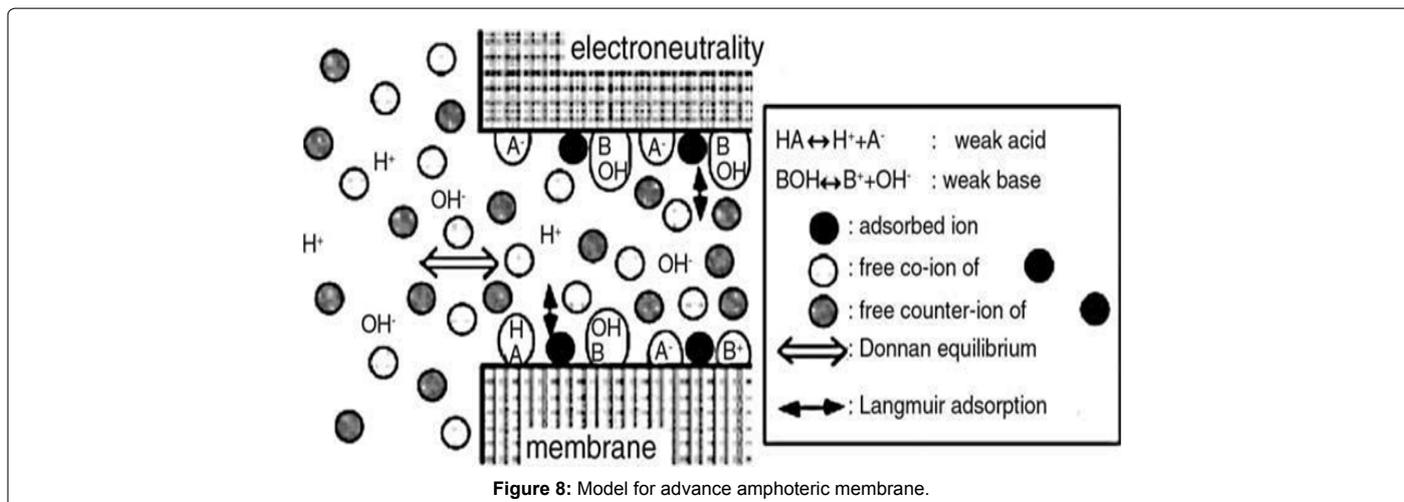


Figure 8: Model for advance amphoteric membrane.

allows passage of cations while restricting passage of anions depends on two parameters:

- (1) Relative affinity of cations with membrane;
- (2) Difference in migration speeds of cations in selective membrane barrier.

For the sake of simplicity an important ratio of given cation to reference cation (This ratio is termed as permeated equivalent) is calculated and usually Na⁺ (sodium) ions. Na⁺ ions are commonly used as reference cation. In order to calculate permselectivity of specified cation, its permeated equivalent value is required.

The required permselectivity between sn⁺ and Na⁺ ions is given by following relation:

$$P(S/Na) = (TS/TNa)/(CS/CNa) \quad (1)$$

TS and TNa are transport numbers symbols of sn⁺ and sodium ions through selective barrier of membrane while cations concentration is represented by symbol C on the membrane outside surface of desalting side solution. If the permeated equivalent and transport number of sn⁺ are represented by Lsi and Tsi,

$$Tsi = Lsi / \sum kLk$$

Then eq. 1 will be modified as,

$$P(S/Na) = (LS \times CNa / LNa \times CS) \quad (2)$$

Lsi can be expressed as the product of mobility of cation sn⁺ and cation sn⁺ concentration in membrane phase. Lsi = Usi × CRsi

Where, Usi (cm²/(de/dy × s)) = mobility of cation; CRsi = concentration of cation. Electrical gradient is depicted by de/dy. On inserting value of Lsi in eq. 2,

$$P(S/Na) = Usi \times CNa \times CRs / UNa \times CRNa \times Cs \quad (3)$$

The relation used for equilibrium constant (KE) between given cation and sodium ions in the membrane phase is given by following equation:

$$KE(S/Na) = CRS / CRNa / CS / Can \quad (4)$$

Upon inserting eq. 4 in eq. 3,

$$P(S/Na) = (Usi / UNa) \times KE(S/Na) \quad (5)$$

In this case effect of diffusional boundary layer at membrane-solution interface on permselectivity will be completely vanished.

The conditions of diffusional boundary layer that exists on the interface of membrane-solution on the desalting side greatly affect permselectivity and it is generally seen that cations permeation rate that selectively permeated reduces while the permeation rate of cations that are permeated reluctantly increases. This all happened because of compositional changes of cations that takes place on layer from that of bulk solution [31].

Membrane fouling of ED system

The existence of certain salts layer on membrane surfaces cause to reduce the membrane recovery rate. Fouling potential increases due to the high concentration of salts in ED systems. Through ion exchange membrane the transport of ions increases with the help of electrical current and with time the membrane selective surface precipitation occurs. With the passage of time accumulation on membrane surface increases and a layer is formed called gel layer. As accumulation charge increases in membrane surface precipitation increase and thinner and denser gel layer formed. Restricted ion movement is due to the formation of the gel layer. Foulant and structure of membrane tell the gel layer concentration and electrical resistance. Current efficiency increases due to increase in current density of ED system. The upper value of current density not achieved due to polarization phenomena. Process efficiency decreases due to the high fouling of high current value than limiting current. Blockages are of three types divalent or trivalent ion hydroxides formed, proteins or colloidal substance accumulation on surface, disruption of membrane.

Anionic membranes fouling are due to presence of organic substance humates, organic substances, and dedokosil sodium benzenesulfonate. Cation membranes blockage in surface and inside is due to accumulation of colloidal materials, proteins, and hydroxyl ions. The fouling degree is affected by the chain length which is the physical feature of fouling molecule. Harder fouling is caused by long chain molecules. Adsorbed rate for the surface of membrane is higher and this is due to the less solubility of large molecules. Other parameter which affects the fouling is current density. Rate of fouling is high when the current density is high. The high hydrodynamics pressure drops and decline in flow is due to the presence of suspended solids, silicates, and less soluble substance which is present in wastewater. Membrane

structures damage physically due to increase to electrical resistance caused by precipitation on the membrane surface.

Increase in concentration caused to increase in membrane fouling if internal membrane fouling control is less. High concentration influences on surface fouling in great. Additional hydraulic resistance is due to protein fouled that can also cause for the formation of gel layer in surface. The waste water treatment of chloralkali industry the presence of calcium and magnesium hydroxide on inside and surface of membrane is reported with the help of studies. Reversible situation achieved due to the precipitation of inorganic salts which yield crystal formation of membrane surface.

Other different compounds which give different precipitation rate of ED process are given below dissolved organic matter, alkaline metals, heavy metals, biomass, organic and inorganic colloids, dissolved gases. Different types of organic solutes effect on the fouling in ED process studies by Lindstrand [32]. Anion and cation selective membrane studied shows that an extraordinary contrast exists between them. A little bit enhancement of cation selective membrane resistance of the meanwhile all solutes fouled the anion selective membranes. This is surely due to the electrostatic attraction between opposite charged organic molecules. Other factor which affects the solubility of organic acids is solution pH, molecular size. Lower solubility, lower mobility, and high adsorption attraction to membrane is enhance the fouling of large acid molecules [1,2,33].

Damages of membrane fouling

The restricted use of ED system is due to fouling reason. Due to fouling the most important parameter which is affected is membrane selectivity which causes to enhance energy demand. Increase in electrical hindrance to anion-exchange membranes is due to precipitation of organic anions. Anion-exchange membranes are more prone to fouling due to the organic compounds.

Important types of pollutants which affect the membrane process are colloidal particles. Stay of physical and chemical factors on the membrane surface which causes by the colloidal substance came with water together with concentration growth rate, salinity, flocculation, surface interaction. Performance of membrane decreased with colloidal foulant many of these types found in literature [33].

Anion exchange membranes are not fouled by the presence of sodium carbonate showed by Araya-Farias and Bazinet [34]. Increase in salt concentration in feed water fouling was observed. Cation exchange membranes contact with sodium chloride the precipitation of calcium hydroxide precipitated studies [35].

In membrane chemical failure and fouling enhance due to destruction in mechanical structure. After a running of 2 years operation fouled membrane decrease in 93% of its toughness and decline in fracture point up to 87% was revealed by Ghalloussi [11]. ED system utilized in food industry fouled by decomposition mechanism no comprehensive study is available even nowadays.

Membrane surface was fouled by polarization process due to the H⁺ ions which is specified by Korngold [36]. Flow velocity reduced due to enhancement of colloid concentration and current density.

Prevention or removal of membrane fouling

To lesson ED membrane fouling much attention is paid. To study this many procedures developed feed solution pre-treatment, process condition optimization, properties of membrane, cell compartments

turbulence. To clean the surface of membrane different methods utilized physical and chemical. Many factors tell which process is utilized for chemical cleaning the most important parameter is type of membrane, intensity of pollution, and type of pollutants. For surface cleaning of membrane, the chemicals recommended by manufacture are high and low strength acids, alkali, complex agents (EDTA), enzymes, detergents and disinfectants (NaCl, H₂O₂). To overcome membrane fouling operation condition is considered further to consider these strategies needed more chemicals and working equipment which enhance the cost of operation [37]. Physical mode of operation for membrane having driving force pressure surface cleaning is done by vibration, air sparge flushing, back washing and CO₂ back permeating. Replacement method to clean membrane surface effectively is ultrasound [38].

Colloids mostly found in natural water are clay, organic deposits, carbon, dust, iron and aluminum hydroxide. To lessen the colloidal pollution chemical cleaning, Pre-filtration, cross flow is utilized. The movement towards the foulants is due to the electrical potential caused by difference in surface charged and electrical conductivity. Increase in resistance is due to the accumulation of the foulants on membrane surface. With the help of alternative electric current, the accumulated foulants removed. Presence of gel layer at ion exchange membrane surface irregularity occurs with the help of pulsed electric fields in ED process [37].

ED membrane fouling due to organic is the presence of humic matters. Disintegration of organic anions on anion exchange membrane as humic acid. Puncture of small organic anions into membrane holding it inside the membrane and caused membrane fouling. Electrical resistance enhances more rapidly due to this reason. Pre-treatment is the necessary step to stop this effect through other detergents. To lesson these affects, different studies recently zeta potential control and pulsed voltage. Membrane structure is damaged by acid-base, mechanical cleaning [39]. Anion exchange membrane surface fouled by sedimentation of negatively charged humate in ED process. ED system sodium chloride solution having 100 mg/L was used. Mobility of charged ions increase in fouling layer due to the pulsed electrical fields and ED cell electrical resistance reduced. Factors on which fouling of membrane depends are feed water specification, membrane characteristic, foulant's physical and chemical properties (Table 3) [37].

Cation exchange membrane affected by fouling in ED process used for treating model solution with composition in calcium and carbonate was reported by Bazinet and Araya-Farias [35]. It was revealed that no fouling happened at concentration of 400 to 800 mg/L of CaCl₂ with no carbonate. Deposition was found only at 400 mg/L CaCl₂ with carbonate. Buffering capacity of the carbonate was given as reason of this difference. System is affected by buffering capacity with and without sodium carbonate [35]. Both researchers in 2006 study found that fouling could be seen on membrane when it was brought in contact with solution of concentration 400, 800 and 1600 mg/L calcium chloride with sodium carbonate. It was depicted that fouling on the surface of anionic membrane exists in two forms, either cubic or crystalline material. The conductivity and thickness of membrane was not drastically affected by fouling and membrane could get rid of fouling by soaking with HCl [35]. Fouling caused by biomass was studied by Bernardes [40]. Microorganisms such as algae, bacteria, fungi in feed water produce biomass and thin layer of fouling is formed on the surface of membrane. Microorganisms could be destroyed by sterilization process. Degradation of polymer membrane by hydrolysis is caused by the biological fouling.

In brackish water desalination industry natural organic matter is

Types of Foulant	Definition	Foulants formula	Charge properties	Cleaning methods
Scale	Precipitates of less soluble salts in the solution	CaCO ₃ , CaSO ₄ ·2H ₂ O, BaSO ₄ , SrSO ₄ , SiO ₂	Non	1. pH adjustment 2. Use of citric acid or EDTA
Colloids	Stack suspended membrane surfaces	SiO ₂ , Fe(OH) ₃ , Al(OH) ₃ , Cr(OH) ₃	Negative	1. Pretreatment with MF, UF, higher flow rate 2. pH adjustment
Organics	Adsorption of organic species to membrane surfaces	Macromolecules, Proteins, whey, polyelectrolytes, humate	Negative	1. Pretreatment with MF, UF, activated carbon 2. Clean with NaOH

Table 3: Varies types of Foulants and cleaning methods used in ED system.

another cause of membrane fouling and in order to get rid of fouling caused by natural organic matter current flow direction is reversed periodically. The method in which current direction is reversed is called EDR and it is similar to ED just polarity of electrodes is reversed. To accomplish EDR completely some pre-treatment is required [41-46].

Conclusion

ED process has been found reliable treatment method for waste water since more than half-a-century. In comparison to other separation techniques such as reverse osmosis, ultra-filtration, nano-filtration and ion exchange, ED has shown more benefits. Some of the advantages are as follows: no need of osmotic pressure, no additional chemical requirement, high quality product, environment friendly. The main demerit of ED process is membrane fouling which causes increase in membrane resistance, flux decline, decreases ions migration yield and high level of polarization.

ED process can help in recycling of reuse valuable products and it has regarded as an important technique for sustainable development because it is an environment friendly process. Its driving force is electric potential difference and thus biomass and organic materials will remain in product stream. EDR could be used to prevent fouling and it is more efficient technique. More research is required for improvement of ED process and for developing new cleaning methods in order to prevent fouling.

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